

Nickel–zinc battery

From Wikipedia, the free encyclopedia

Jump to: [navigation](#), [search](#)

Nickel–zinc battery	
Specific energy	100 W·h/kg
Energy density	280 W·h/L
Specific power	> 3000 W/kg
Energy/consumer-price	2–3 Wh/US\$
Nominal cell voltage	1.65 V

A **nickel–zinc battery**, abbreviated **NiZn**, is a type of [rechargeable battery](#) that may be used in cordless power tools, cordless telephones, digital cameras, battery operated lawn and garden tools, professional photography, [flashlights](#), electric bikes, and light [electric vehicle](#) sectors, among other uses.

Larger [nickel–zinc](#) battery systems have been known for over 100 years. Since 2000, development of a stabilized zinc [electrode](#) system has made this technology viable and competitive with other commercially available rechargeable battery systems.

Contents

[\[hide\]](#)

- [1 History](#)
- [2 Applications](#)
- [3 Electrochemistry](#)
- [4 Battery life](#)
- [5 Advantages](#)
- [6 Disadvantages](#)
 - [6.1 Cost](#)
- [7 Charging](#)
- [8 See also](#)
- [9 References](#)

History[[edit](#)]

In 1901, [Thomas Alva Edison](#) was awarded [U.S. Patent 684,204](#) for a rechargeable nickel–zinc battery system.^[1]

The battery was later developed by the [Irish](#) chemist Dr. James J. Drumm (1897–1974),^[2] and installed in four two-car [Drumm railcar](#) sets between 1932 and 1948 for use on the [Dublin–Bray](#) railway line. Although successful, they were withdrawn when the batteries wore out. Early nickel–zinc batteries were plagued by a limited number of discharge/recharge cycles. In the 1960s, nickel–zinc batteries were investigated as an alternative to silver–zinc batteries for military applications, and in the 1970s, were again of interest for electric vehicles.^[3] A company called *Evercel Inc.* developed and patented several improvements in nickel–zinc batteries, but withdrew from that area in 2004.^[4]

Applications[[edit](#)]

Nickel–zinc batteries have a charge/discharge curve similar to 1.2 V [NiCd](#) or [NiMH](#) cells, but with a higher 1.6 V nominal voltage.^[5]

Nickel–zinc batteries perform well in high-drain applications, and may have the potential to replace [lead–acid batteries](#) because of their higher energy-to-mass ratio and higher power-to-mass ratio (up to 75% lighter for the same power).^[6] NiZn are cheaper than [nickel-cadmium batteries](#),^[6] and are expected to be priced somewhere in between NiCd and lead–acid types. NiZn may be used as a substitute for [nickel-cadmium](#). The European Parliament has supported bans on cadmium-based batteries,^[1] and nickel–zinc offers the European power tool industry a good alternative.

After about 30 cycles, the self-discharge rate may increase, so after charging the batteries may not stay "fresh" as long. For applications requiring high-power and high-voltage, NiZn is a good battery choice when good charge-retention is not required.^[7]

However, if used properly, NiZn batteries can last for hundreds of cycles.^[8]

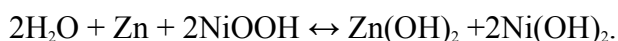
Electrochemistry[[edit](#)]



This section **does not cite any sources**. Please help improve this section by [adding citations to reliable sources](#). Unsourced material may be challenged and [removed](#). *(July 2015)*

Water is consumed and generated on the discharge and charge cycles respectively.

Discharge Reaction is left to right:



Electrochemical open circuit voltage potential: ~1.73 V.

Battery life[[edit](#)]

Compared with cadmium hydroxide, the tendency of zinc hydroxide to dissolve into solution and not fully migrate back to the cathode during recharging has, in the past, presented challenges for the commercial viability of the NiZn battery.^{[1][3]} The zinc's reluctance to fully return to the solid electrode adversely manifests itself as shape change and [dendrites](#) (or "[whiskers](#)"), which may reduce the cell discharging performance or, eventually, short out the cell, resulting in a low cycle life.

Recent advances have enabled manufacturers to greatly reduce this problem. These advances include improvements in electrode separator materials, inclusion of zinc material stabilizers, and electrolyte improvements (i.e. by using [phosphates](#)). One manufacturer, ([PowerGenix](#)), which has developed 1.6V batteries, has claimed battery cycle life comparable to NiCd batteries.^[9]

Battery cycle life is most commonly specified at a discharge depth of 80 percent of rated capacity and assuming a one-hour discharge current rate. If the discharge current rate is reduced, or if the depth of discharge is reduced, then the number of charge-discharge cycles for a battery increases. When comparing NiZn to other battery technologies, cycle life specifications may vary with other battery technologies, depending on the discharge rate and depth of discharge that were used.

Advantages[[edit](#)]

Nickel–zinc cells have an open circuit voltage of 1.85 volts when fully charged,^[10] and a nominal voltage of 1.65 V. This makes NiZn an excellent replacement for electronic products that were designed to use alkaline primary cells (1.5 V). NiCd and NiMH both have nominal cell voltages of 1.2 V, which may cause some electronic equipment to shut off prior to a complete discharge of the battery because the minimal operating voltage is not provided.

Newer, more powerful cells with up to 800 cycles/life can be an alternative to Li-Ion batteries for electric vehicles. Due to their higher voltage, fewer cells are required (compared to NiCd and NiMH) to achieve a given battery-pack voltage, reducing pack weight and size and improving pack reliability. They also have low internal [impedance](#) (typically 5 [milliohms](#)), which allows for high battery discharge rates, up to 50C. (C is battery capacity in Ah, divided by one hour.)^[citation needed]

NiZn batteries do not use mercury, lead, or cadmium, or [metal hydrides](#) that are difficult to recycle.^[11] Both nickel and zinc are commonly occurring elements in nature, and can be fully recycled.

NiZn cells use no flammable active materials or organic electrolytes, and the newest models use polymeric separators to resolve the dendrites problem.

Properly designed NiZn cells can have very high power density and low temperature discharging performance, and also can be discharged to 100% and recharged without problems. They are now available in sizes up to F and 50Ah/ prismatic cell.

Zinc is cheap and abundant metal (the 24th most abundant element in the Earth's crust) and it is not dangerous to health. Common oxidation is +2 so charge and discharge move two electrons instead of one as in NiMH batteries.

Disadvantages[[edit](#)]



This section **does not cite any sources**. Please help improve this section by [adding citations to reliable sources](#). Unsourced material may be challenged and [removed](#). *(July 2015)*

NiZn batteries cannot be discharged to 0 V, end-of-discharge voltage must be 1.3 V or the cell will be damaged.

Cost[[edit](#)]

Current Nickel Zinc batteries use nickel foam and copper foam current collectors for active materials; they are expensive, but new cheaper carbon-based current collectors are under development. Moreover, special rechargers must be used for NiZn batteries, rechargers for NiCd or NiMH are not suitable and will destroy the NiZn cell.

Charging[[edit](#)]

NiZn technology is well suited for fast recharge cycling, as optimum charge rates of [C](#) or [C/2](#) are preferred.^[12]

Known charging regimes include a constant current of C or C/2 to cell voltage = 1.9 V. Maximum charge time is given variously as 2½ hours^[12] and 3 hours.^[13] [Trickle charging](#) is not recommended, as recombination is not provided for, and excess [hydrogen](#) will eventually vent, adversely affecting battery cycle life.

See also[[edit](#)]

- [List of battery types](#)
- [List of battery sizes](#)
- [Comparison of battery types](#)
- [Nickel–cadmium battery](#)
- [Zinc–air battery](#)

References[[edit](#)]

1. ^a ^b ^c ^e ["Building A Better Battery"](#), Kerry A. Dolan, Forbes.com, *Forbes* magazine, 11 May 2009, accessed 2011-02-12, [Forbes-44](#).
2. [Jump up](#) ^a ^b ^c ^e ["Famous Irish Chemists: James J. Drumm"](#). *Ul.ie*. Retrieved 2012-07-01.
3. ^a ^b ^c ^e [David Linden \(ed\)., *Handbook of Batteries*, McGraw Hill, 2002, ISBN 0-07-135978-8, chapter 31.](#)
4. [Jump up](#) ^a ^b ^c ^e [Evercel financial statement 2007](#), Evercel.com, page 9, retrieved 23 November 2010.
5. [Jump up](#) ^a ^b ^c ^e [Battery-meter-problem](#), NiZn discharge curves and camera voltage cutoffs, PentaxForums.com
6. ^a ^b ^c ^e ["Nickel Zinc"](#). *EnerSys.com*. [EnerSys](#). Retrieved 13 July 2015.
7. [Jump up](#) ^a ^b ^c ^e ["A Review Of NiZn Batteries"](#). 16 March 2012.
8. [Jump up](#) ^a ^b ^c ^e [Thomas, Justin \(16 March 2012\). "A Review Of NiZn Batteries"](#). *MetaEfficient.com*. inSync Theme. Retrieved 13 July 2015.
9. [Jump up](#) ^a ^b ^c ^e ["A Brief History of Battery Developments"](#), PowerGenix.com, 2010, retrieved 12 February 2011.
10. [Jump up](#) ^a ^b ^c ^e [\[1\] New NiZn batteries offer lightning-fast recycle](#)
11. [Jump up](#) ^a ^b ^c ^e ["PowerGenix NiZn Material Safety Data Sheet"](#), PowerGenix.com, accessed=2011-02-12.
12. ^a ^b ^c ^e ["PowerGenix AA Battery Specifications" \(PDF\)](#). Retrieved 2012-07-01.
13. [Jump up](#) ^a ^b ^c ^e ["PowerGenix NiZn Quick Charger"](#). *Powergenix.com*. Retrieved 2012-07-01.

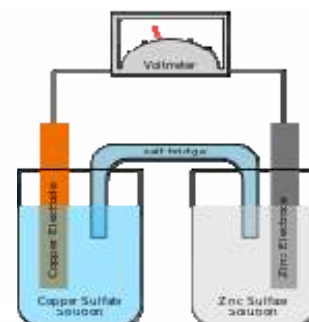
[hide]

- [v](#)
- [t](#)
- [e](#)

Galvanic cells

Types

- [Voltaic pile](#)
- [Battery](#)
 - [Flow battery](#)
 - [Trough battery](#)
- [Concentration cell](#)
- [Fuel cell](#)
- [Thermogalvanic cell](#)



Primary cell

(non-rechargeable)

- [Alkaline](#)
- [Aluminium-air](#)

- [Bunsen](#)
- [Chromic acid](#)
- [Clark](#)
- [Daniell](#)
- [Dry](#)
- [Edison-Lalande](#)
- [Grove](#)
- [Leclanché](#)
- [Lithium](#)
- [Mercury](#)
- [Nickel oxyhydroxide](#)
- [Silicon-air](#)
- [Silver oxide](#)
- [Weston](#)
- [Zamboni](#)
- [Zinc-air](#)
- [Zinc-carbon](#)

[Secondary cell](#)
(rechargeable)

- [Automotive](#)
- [Lead-acid](#)
 - [gel / VRLA](#)
- [Lithium-air](#)
- [Lithium-ion](#)
 - [Lithium polymer](#)
 - [Lithium iron phosphate](#)
 - [Lithium titanate](#)
 - [Lithium-sulfur](#)
 - [Dual carbon battery](#)
- [Molten salt](#)
- [Nanopore](#)
- [Nanowire](#)
- [Nickel-cadmium](#)
- [Nickel-hydrogen](#)
- [Nickel-iron](#)
- [Nickel-lithium](#)
- [Nickel-metal hydride](#)
- **Nickel-zinc**
- [Polysulfide bromide](#)

	<ul style="list-style-type: none"> Zinc–cerium
Cell parts	<ul style="list-style-type: none"> Anode Binder Catalyst Cathode Electrode Electrolyte Half-cell Ions Salt bridge Semipermeable membrane

Retrieved from "https://en.wikipedia.org/w/index.php?title=Nickel-zinc_battery&oldid=707700996"

Categories:

- [Nickel](#)
- [Rechargeable batteries](#)
- [Zinc](#)

Hidden categories:

- [Use dmy dates from January 2012](#)
- [Articles needing additional references from July 2015](#)
- [All articles needing additional references](#)
- [All articles with unsourced statements](#)
- [Articles with unsourced statements from January 2015](#)

Navigation menu

Personal tools

- Not logged in
- [Talk](#)
- [Contributions](#)
- [Create account](#)
- [Log in](#)

Namespaces

- [Article](#)

- [Talk](#)

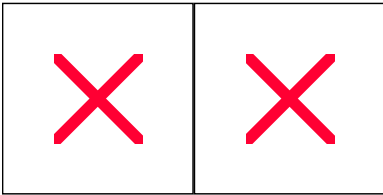
Variants

Views

- [Read](#)
- [Edit](#)
- [View history](#)

More

Search



Navigation

- [Main page](#)
- [Contents](#)
- [Featured content](#)
- [Current events](#)
- [Random article](#)
- [Donate to Wikipedia](#)
- [Wikipedia store](#)

Interaction

- [Help](#)
- [About Wikipedia](#)
- [Community portal](#)
- [Recent changes](#)
- [Contact page](#)

Tools

- [What links here](#)
- [Related changes](#)
- [Upload file](#)
- [Special pages](#)

- [Permanent link](#)
- [Page information](#)
- [Wikidata item](#)
- [Cite this page](#)

Print/export

- [Create a book](#)
- [Download as PDF](#)
- [Printable version](#)

Languages

- [Deutsch](#)
- [Español](#)
- [فارسی](#)
- [Français](#)
- [Italiano](#)
- [Română](#)
- [Русский](#)

[Edit links](#)

- This page was last modified on 1 March 2016, at 08:36.
- Text is available under the [Creative Commons Attribution-ShareAlike License](#); additional terms may apply. By using this site, you agree to the [Terms of Use](#) and [Privacy Policy](#). Wikipedia® is a registered trademark of the [Wikimedia Foundation, Inc.](#), a non-profit organization.
- [Privacy policy](#)
- [About Wikipedia](#)
- [Disclaimers](#)
- [Contact Wikipedia](#)
- [Developers](#)
- [Cookie statement](#)
- [Mobile view](#)
-