Basic Chemical Material Manufacture:

1: Hydrogen & Oxygen can be obtained in large quantities from water H2O by electrolysis.

2: Hydrogen - H2 & Nitrogen -N2 can be used to make Ammonia NH3 & Nitrates – both useful as a necessary supplementary agricultural fertilizer.

Sodium-nitrate [Chile Saltpeter]- $NaNO_3$
Potassium-nitrate - KNO_3
Ammonium-nitrate - $NH4NO_3$

3: With access to high-temperature steam and charcoal – Carbon-monoxide can be produced:

Carbon C + Water H2O + heat ----------> Carbon-monoxide CO + Hydrogen H2

4: Using certain catalysts and Hydrogen - Carbon-monoxide can be transformed into Methane CH4:

Carbon-monoxide CO + Hydrogen 3H2 + heat ----------> Methane CH4 + Water H2O

5: Carbon-monoxide can also be transformed into liquid alcohol – Methanol:

Carbon-monoxide CO + Hydrogen 2H2 + heat => Methanol CH3OH

Methanol can be used as fuel in spark-ignition engines and as a solvent.

6: Methane can likewise be used as fuel, a solvent and by the use of other catalysts transformed into Acetylene – a basis raw-material for a number of organic materials.

Methane 2CH4 + heat -------> Acetylene CH ≡ AcetyleneCH + Hydrogen 2H2

The necessary process heat can be provided from wood-powered thermal/steam electric generation sources.
However it may be noted that the original start-material used is charcoal - if coke [from coal] is available this can be used as a raw-material instead of charcoal.

Even though adequate electric power is available – the limiting factor of such chemical industrial-process will be the availability of charcoal or other carbon fuel.

Such a bottle-neck can be overcome - using atmospheric Carbon–dioxide CO2 - provided large quantities of power are available

Carbon-dioxide is first prepared as in dry-ice manufacture [cold-solid CO2]. Thereafter fine particles of Zinc react under pressure with the Carbon-dioxide.

CO2 + Zn => CO + ZnO

The resulting Carbon-monoxide can be hydrogenerated [disolved under a hydration process] with Hydrogen to Methane or Methanol. The remaining Zinc-oxide can be regenerated under high pressure & high temperature with Hydrogen to Zinc. The Zinc is not consumed but changes between metallic Zinc and Zinc-oxide.

It is also possible to produce Methane & Methanol by direct reaction between atmospheric Carbon-dioxide and Hydrogen produced by electrolysis of water.

$CO_2 + 4H_2 \dashrightarrow CH_4 + 2H_2O$

However in a technical production this process requires large and stabile pressure-tanks,

This process is similar to photosynthesis in green plants – whereby Carbohydrates/starch are formed following the formula:

$nCO_2 + nH_2O \xrightarrow{\text{\& sunlight \&}} [CH_2O] + nO_2$

This process is in two stages as follows:

1: $2H_2O + 2X \xrightarrow{\text{\& sunlight}} \xrightarrow{\text{\& sunlight}} 2XH_2 + O_2$
2: $CO_2 + 2XH_2 \rightarrow CH_2O + H_2O$

This second stage is the real hydrogenerating process of Carbon-dioxide. However it is in this first stage that the splitting of water by Chlorophyll and sunlight;– That is the energy-demanding process of the synthesis of Carbohydrate by plants occurs.

Extra notes:

Carbon-dioxide CO_2 can be used as a raw-material to produce organic-chemical compounds with Zinc as a catalyst.
$CO_2 + Zn \dashrightarrow CO + ZnO$
$ZnO + H_2 \dashrightarrow Zn + H_2O$
Result: $CO_2 + H_2 \dashrightarrow CO + H_2O$

With CO => conventional synthesis of Methane ----------> Acethylene.
Methanol ----------> Formaldehyde CH_2O

Hydrogen used in this above process can be obtained by the electrolysis of water.

A possible hydrogen process under high pressure can be as follows:
$CO_2 + 3H_2 \dashrightarrow CH_3OH + H_2O$

Iron-ore:
Iron-oxide ore can be reduced by Hydrogen obtained by electrolysis.
$Fe_2O_3 + 3H_2 \dashrightarrow 2Fe + 3H_2O$

Basic manufacture of lime for construction cement & for magnesium as a substitute for copper for electrical windings & power lines etc – for use during periods of very severe material shortages:

Cement:

Burn limestone to unslaked lime or quick lime & carbon dioxide
$CaCO3 + CO2$

Magnesium:

Lime CaO is obtained by burning of limestone CaCO3. The Carbon-dioxide CO2 disappears leaving lime.

A small bay or inlet on the sea coast containing sea-water is enclosed by a constructed dam.

Lime is then continuously poured into this sea-salt water. Magnesium-salt in solution precipitates & falls to the bottom as Magnesium-hydroxide.

This is then collected by shovel & thereafter treated in a thick semi-liquid form with hydrochloric acid - HCl.

Hydrochloric acid is obtained by electrolysis of concentrated sea-water to Chlorine Cl + Hydrogen => HCl

Hydrochloric acid and Magnesium-hydroxide mixed together will give Magnesium-chloride.

This is first dried and then is ready for smelting by electrolysis by 10-12 VDC at several thousand amperes. The Magnesium will be deposited on the cathode as melted metal.

Electrical power is required at two stages in the process:
A: Electrolysis of the sea-water
B: Electrolysis of Magnesium-chloride

Heat - possibly from wood - will be required at three stages:
A: Preparation of Lime from limestone.
B: Concentration of sea-water
C: Concentration of the Magnesium-chloride solution.

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