

Physical chemistry

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Physical chemistry is the study of macroscopic, atomic, subatomic, and particulate phenomena in chemical systems in terms of the principles, practices and concepts of physics such as motion, energy, force, time, thermodynamics, quantum chemistry, statistical mechanics, analytical dynamics and chemical equilibrium.

Physical chemistry, in contrast to chemical physics, is predominantly (but not always) a macroscopic or supra-molecular science, as the majority of the principles on which it was founded relate to the bulk rather than the molecular/atomic structure alone (for example, chemical equilibrium and colloids).

Some of the relationships that physical chemistry strives to resolve include the effects of:

1. Intermolecular forces that act upon the physical properties of materials (plasticity, tensile strength, surface tension in liquids).
2. Reaction kinetics on the rate of a reaction.
3. The identity of ions and the electrical conductivity of materials.
4. Surface chemistry and electrochemistry of cell membranes.^[1]
5. Interaction of one body with another in terms of quantities of heat and work called thermodynamics.
6. Transfer of heat between a chemical system and its surroundings during change of phase or chemical reaction taking place called thermochemistry
7. Study of colligative properties of number of species present in solution.
8. Number of phases, number of components and degree of freedom (or variance) can be correlated with one another with help of phase rule.
9. Reactions of electrochemical cells.

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Key concepts

The key concepts of physical chemistry are the ways in which pure physics is applied to chemical problems.

One of the key concepts in classical chemistry is that all chemical compounds can be described as groups of atoms bonded together and chemical reactions can be described as the making and breaking of those bonds. Predicting the properties of chemical compounds from a description of atoms and how they bond is one of the major goals of physical chemistry. To describe the atoms and bonds precisely, it is necessary to know both where the nuclei of the atoms are, and how electrons are distributed around them.^[2]

Quantum chemistry, a subfield of physical chemistry especially concerned with the application of quantum mechanics to chemical problems, provides tools to determine how strong and what shape bonds are,^[2] how nuclei move, and how light can be absorbed or emitted by a chemical compound.^[3] Spectroscopy is the related sub-discipline of physical chemistry which is specifically concerned with the interaction of electromagnetic radiation with matter.

Another set of important questions in chemistry concerns what kind of reactions can happen spontaneously and which properties are possible for a given chemical mixture. This is studied in chemical thermodynamics, which sets limits on quantities like how far a reaction can proceed, or how much energy can be converted into work in an internal combustion engine, and which provides links between properties like the thermal expansion coefficient and rate of change of entropy with pressure for a gas or a liquid.^[4] It can frequently be used to assess whether a reactor or engine design is feasible, or to check the validity of experimental data. To a limited extent, quasi-equilibrium and non-equilibrium thermodynamics can describe irreversible changes.^[5] However, classical thermodynamics is mostly concerned with systems in equilibrium and reversible changes and not what actually does happen, or how fast, away from equilibrium.

Which reactions do occur and how fast is the subject of chemical kinetics, another branch of physical chemistry. A key idea in chemical kinetics is that for reactants to react and form products, most chemical species must go through transition states which are higher in energy than either the reactants or the products and serve as a barrier to reaction.^[6] In general, the higher the barrier, the slower the reaction. A second is that most chemical reactions occur as a sequence of elementary reactions,^[7] each with its own transition state. Key questions in kinetics include how the rate of reaction depends on temperature and on the concentrations of reactants and catalysts in the reaction mixture, as well as how catalysts and reaction conditions can be engineered to optimize the reaction rate.

The fact that how fast reactions occur can often be specified with just a few concentrations and a temperature, instead of needing to know all the positions and speeds of every molecule in a mixture, is a special case of another key concept in physical chemistry, which is that to the extent an engineer needs to know, everything going on in a mixture of very large numbers (perhaps of the order of the Avogadro constant, 6×10^{23}) of particles can often be described by just a few variables like pressure, temperature, and concentration. The precise reasons for this are described in statistical mechanics,^[8] a specialty within physical chemistry which is also shared with physics. Statistical mechanics also provides ways to predict the properties we see in everyday life from molecular properties without relying on empirical correlations based on chemical similarities.^[5]

History

The term "physical chemistry" was coined by Mikhail Lomonosov in 1752, when he presented a lecture course entitled "A Course in True Physical Chemistry" (Russian: «Курс истинной физической химии») before the students of Petersburg University.^[9] In the preamble to these lectures he gives definition: "Physical chemistry is the science that must explain under provisions of physical experiments the reason for what is happening in complex bodies through chemical operations".

Modern physical chemistry originated in the 1860s to 1880s with work on chemical thermodynamics, electrolytes in solutions, chemical kinetics and other subjects. One milestone was the publication in 1876 by Josiah Willard Gibbs of his paper, *On the Equilibrium of Heterogeneous Substances*. This paper introduced several of the cornerstones of physical chemistry, such as Gibbs energy, chemical potentials, and Gibbs' phase rule.^[10] Other milestones include the subsequent naming and accreditation of enthalpy to Heike Kamerlingh Onnes and to macromolecular processes.

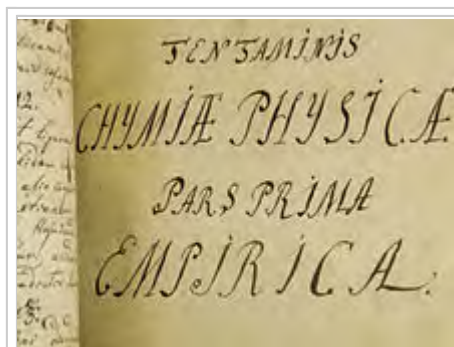
The first scientific journal specifically in the field of physical chemistry was the German journal, *Zeitschrift für Physikalische Chemie*, founded in 1887 by Wilhelm Ostwald and Jacobus Henricus van 't Hoff. Together with Svante August Arrhenius,^[11] these were the leading figures in physical chemistry in the late 19th century and early 20th century. All three were awarded with the Nobel Prize in Chemistry between 1901-1909.

Developments in the following decades include the application of statistical mechanics to chemical systems and work on colloids and surface chemistry, where Irving Langmuir made many contributions. Another important step was the development of quantum mechanics into quantum chemistry from the 1930s, where Linus Pauling was one of the leading names. Theoretical developments have gone hand in hand with developments in experimental methods, where the use of different forms of spectroscopy, such as infrared spectroscopy, microwave spectroscopy, EPR spectroscopy and NMR spectroscopy, is probably the most important 20th century development.

Further development in physical chemistry may be attributed to discoveries in nuclear chemistry, especially in isotope separation (before and during World War II), more recent discoveries in astrochemistry,^[12] as well as the development of calculation algorithms in the field of "additive physicochemical properties" (practically all physicochemical properties, such as boiling point, critical point, surface tension, vapor pressure, etc.—more than 20 in all—can be precisely calculated from chemical structure alone, even if the chemical molecule remains unsynthesized), and herein is concentrated practical importance of contemporary physical chemistry.

See Group contribution method, Lydersen method, Joback method, Benson group increment theory, QSPR, QSAR

Journals



Fragment of M. Lomonosov's manuscript 'Physical Chemistry' (1752)

Some journals that deal with physical chemistry include:

- Zeitschrift für Physikalische Chemie (1887)
- Journal of Physical Chemistry A (from 1896 as *Journal of Physical Chemistry*, renamed in 1997)
- Physical Chemistry Chemical Physics (from 1999, formerly Faraday Transactions with a history dating back to 1905)
- Macromolecular Chemistry and Physics (1947)
- Annual Review of Physical Chemistry (1950)
- Molecular Physics (1957)
- Journal of Physical Organic Chemistry (1988)
- Journal of Physical Chemistry B (1997)
- ChemPhysChem (2000)
- Journal of Physical Chemistry C (2007)
- Journal of Physical Chemistry Letters (from 2010, combined letters previously published in the separate journals)

Historical journals that covered both chemistry and physics include *Annales de chimie et de physique* (started in 1789, published under the name given here from 1815–1914).

Branches and related topics

- Thermochemistry
- Chemical kinetics
- Quantum chemistry
- Electrochemistry
- Photochemistry
- Surface chemistry
- Solid-state chemistry
- Spectroscopy
- Biophysical chemistry
- Materials science
- Physical organic chemistry
- Micromeritics

See also

- List of important publications in chemistry#Physical chemistry
- List of unsolved problems in chemistry#Physical chemistry problems
- Physical biochemistry
- Category:Physical chemists

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11. Laidler, Keith (1993). *The World of Physical Chemistry*. Oxford: Oxford University Press. p. 48. ISBN 0-19-855919-4.
12. Herbst, Eric (May 12, 2005). "Chemistry of Star-Forming Regions". *Journal of Physical Chemistry A*. **109** (18): 4017–4029. doi:10.1021/jp050461c. PMID 16833724.

External links

- The World of Physical Chemistry (https://books.google.com/books?id=01LRIPbH80cC) (Keith J. Laidler, 1993)
- Physical Chemistry from Ostwald to Pauling (https://books.google.com/books?id=1UZjU2WfLAoC) (John W. Servos, 1996)
- 100 Years of Physical Chemistry (http://www.amazon.com/Years-Physical-Chemistry-Royal-Society/dp/0854049878/) (Royal Society of Chemistry, 2004)
- Physical Chemistry: neither Fish nor Fowl? (http://www.joachimschummer.net/jslit/physchem.htm) (Joachim Schummer, *The Autonomy of Chemistry*, Würzburg, Königshausen & Neumann, 1998, pp. 135–148)
- Cathedrals of Science (http://www.oup.com/us/catalog/general/subject/Chemistry/?view=usa&ci=9780195321340) (Patrick Coffey, 2008)
- The Cambridge History of Science: The modern physical and mathematical sciences (https://books.google.com/books?id=B3WvWhJTTX8C) (Mary Jo Nye, 2003)



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